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6.2 Define the term chemical equilibrium. At equilibrium the concentration of the NO_2 (g) is 0.2 mol dm^{-3} . The equilibrium constant for the reaction is 171 at 25°C . Calculate the initial number of moles of NO_2 (g) placed in the gas syringe. 6.3 The diagram below shows the reaction mixture in the gas syringe after

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The N_2O_4 gas decomposes to NO_2 gas and the reaction reaches equilibrium according to the following balanced equation: $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $\Delta H < 0$. The graph below shows how the concentrations of the two gases change as a result of changes made to the reaction conditions. N_2O_4 (g) NO_2 (g) 6.1 6.2

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5 kg the 5 kg block hangs vertically as shown in the diagram below a 15 n force is applied to the 8 kg block at an angle of 30° to the horizontal causing the block to slide to the left 15 n 30° 8 kg 5 kg the coefficient of kinetic friction between the 8 kg block and the surface of the table is 0 25

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the net force acting on an object is equal to the rate of change of momentum of the object in the direction of the force 2 or 0

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