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14 1 introduction to oxidation and reduction chemistry Dec 17 2023 on december 15 1967 the 39 year old bridge collapsed killing 46 people the ultimate cause of the collapse was determined to be corrosion of a suspension chain on the ohio side of the bridge corrosion is an example of the type of chemical reaction discussed in this chapter

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simultaneously the term redox comes from the combination of the two processes involved oxidation states the oxidation state of an ion refers to the charge that the ion possess

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what is reduction answer reduction is a process in which electrons are gained oxygen is removed or hydrogen is added during a reaction by a molecule atom or ion example $\text{Zn} + 2\text{H}^+\text{aq} \rightarrow \text{Zn}^{2+}\text{aq} + \text{H}_2\text{g}$ q15 identify the species being oxidized and reduced in the following reaction $3\text{Hg} + 2\text{Fe} + \text{S} \rightarrow 3\text{Hg}_2 + 2\text{FeS}$

oxidation reduction chemistry a level revision Jul 12 2023 the atom that loses electrons is oxidised and the atom that gains electrons is reduced oxidation is an increase in oxidation number reduction is a decrease in oxidation number an oxidising agent removes electrons a reducing agent donates electrons

detailed notes topic 8 redox chemistry and groups 1 2 and Jun 11 2023

topic 8a redox chemistry oxidation and reduction oxidation is the loss of electrons reduction is the gain of electrons this redox rule is remembered using the acronym oilrig oxidation is loss reduction is gain oxidation number oxidation number gives the oxidation state of an element or ionic substance

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copper ii ions gain electrons and gain of electrons is reduction redox is when oxidation and reduction occur simultaneously in the same reaction

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reduction involves a half reaction in which a chemical species decreases its oxidation number usually by gaining electrons the other half of the reaction involves oxidation in which electrons are lost together reduction and oxidation form redox reactions reduction oxidation redox

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