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figure pageindex 1 the structure of the atom atoms have protons and neutrons in the center making the nucleus while the electrons orbit the nucleus the modern atomic theory states that atoms of one element are the same while atoms of different elements are different what three particles make up an atom choose 1 answer protons neutrons and photons a protons neutrons and photons positrons neutrons and electrons b positrons neutrons and electrons protons electrons and positrons c protons electrons and positrons protons neutrons and electrons d protons neutrons and electrons answer atomic number z the atomic number of an element is equal to the number of protons present inside the nucleus of its atoms since an isolated atom has no net charge on it in neutral atoms the total number of electrons is equal to its atomic number the atomic number tells you the number of in one atom of an element it also tells you the number of in a neutral atom of that element the atomic number gives the identity of an element as well as its location on the periodic table the atomic number is always written on the bottom left side of the element the charge if there is one is written on the top right side see structure packet if no charge is given assume that atom is neutral equal protons and electrons what is the mass number of an atom with 3 protons 4 neutrons and 3 electrons how many neutrons are in the nucleus of an atom that has an atomic mass of 36 and an atomic number of 25 quiz yourself with questions and answers for atomic structure practice test so you can be ready for test day explore quizzes and practice tests created by teachers and students or create one from your course material atomic structure and periodic properties 1 how many orbitals are there in a shell of principal quantum number n hint begin with n 1 2 and 3 and see if you can recognize the pattern atomic basics answer key part a atomic structure 1 draw five protons in the nucleus of the atom label them with their charge 2 draw six neutrons in the nucleus of the atom 3 draw two electrons in the first energy level and label them with their charge 4 draw three electrons in the second energy level and label them with their charge 5 the atomic number gives the identity of an element as well as its location on the periodic table no two different elements will have the atomic number the of an element is the average mass of an element s naturally occurring atom or isotopes taking into account the of each isotope atomic structure practice test answer key contains practice tests to support learning in chemistry learning about atoms course chemistry chem253 628documents students shared 628 documents in this course university harvard university academic year 2023-10-01 1/7 respiratory technology procedure and equipment manual

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figure pageindex 1 the structure of the atom atoms have protons and neutrons in the center making the nucleus while the electrons orbit the nucleus the modern atomic theory states that atoms of one element are the same while atoms of different elements are different

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what three particles make up an atom choose 1 answer protons neutrons and photons a protons neutrons and photons positrons neutrons and electrons b positrons neutrons and electrons protons electrons and positrons c protons electrons and positrons protons neutrons and electrons d protons neutrons and electrons

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answer atomic number z the atomic number of an element is equal to the number of protons present inside the nucleus of its atoms since an isolated atom has no net charge on it in neutral atoms the total number of electrons is equal to its atomic number

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the atomic number tells you the number of in one atom of an element it also tells you the number of in a neutral atom of that element the atomic number gives the identity of an element as well as its location on the periodic table

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the atomic number is always written on the bottom left side of the element the charge if there is one is written on the top right side see structure packet if no charge is given assume that atom is neutral equal protons and electrons

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the atomic number gives the identity of an element as well as its location on the periodic table no two different elements will have the atomic number the of an element is the average mass of an element s naturally occurring atom or isotopes taking into account the of each isotope

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the answers labeled i below received two points ii received one point a two points the radii of the alkali metal ions increase with increasing atomic number because i the principle quantum number or shell or energy level increases ii there is an increase in shielding or the number of orbitals increases b two points

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this document contains 11 past paper questions about atomic structure the questions cover topics such as isotopes electronic configuration subatomic particles the periodic table and relative atomic mass

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