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stoichiometry volume volume problems 1 n<sub>2</sub> name what volume of hydrogen is necessary to react with five liters of nitrogen to produce ammonia assume constant temperature and pressure 5 l 3 l 2 what volume of ammonia is produced in the reaction in problem 1 3 502 3 c02 l co

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application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and

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to determine the amounts of product either grams or moles you must start with the limiting reagent use the amount that you have not the amount you need to determine the grams of excess reagent subtract the amount you need from the amount that you have then using the molar mass convert the moles left to grams

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worksheet calculating empirical molecular formulas 1 the empirical formula for the compound having the formula h<sub>2</sub>c<sub>2</sub>o<sub>4</sub> is a c<sub>2</sub>h<sub>2</sub> b co<sub>2</sub>h c coh d c<sub>2</sub>o<sub>4</sub>h<sub>2</sub> e coh<sub>2</sub> 2 calculate the empirical formula of a compound that is 85 6 c and 14 4 h by mass a ch<sub>2</sub> b ch c c<sub>3</sub>h<sub>5</sub> d c<sub>2</sub>h<sub>4</sub> e c<sub>2</sub>h<sub>3</sub>

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this is a comprehensive end of chapter set of practice problems on stoichiometry that covers balancing chemical equations mole ratio calculations limiting reactants and percent yield concepts the links to the corresponding topics are given below

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it is prepared by the reaction of pure sand  $\text{SiO}_2$  with carbon at high temperature carbon monoxide  $\text{CO}$  is the other product of this reaction write the balanced equation for the reaction and calculate how much  $\text{SiO}_2$  is required to produce 3 00 kg of  $\text{SiC}$  answer

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write a balanced molecular equation describing each of the following chemical reactions solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas gaseous butane  $\text{C}_4\text{H}_{10}$  reacts with

diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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stoichiometry to formulate and solve material and energy balance problems in processes with and without chemical reactions with the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams it also elaborates on

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if the molecular weight of glucose is 180 grams mol and the molar mass of ethanol is 46 g mol how many grams of carbon dioxide are produced when 1 mol of glucose is digested via respiration choose 1 answer 88 grams

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