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answer atomic number 50 mass number 118 when referring to an atom we simply use the element s name the term sodium refers to the element as well as an atom of sodium what three particles make up an atom choose 1 answer protons neutrons and photons a protons neutrons and photons positrons neutrons and electrons b positrons neutrons and electrons protons electrons and positrons c protons electrons and positrons protons neutrons and electrons d protons neutrons and electrons answer atomic number z the atomic number of an element is equal to the number of protons present inside the nucleus of its atoms since an isolated atom has no net charge on it in neutral atoms the total number of electrons is equal to its atomic number study with quizlet and memorize flashcards containing terms like the 3 particles of the atom are their respective charges are the number of protons in one atom of an element determines the atom s and the number of electrons determines the of the element and more 19 multiple choice questions definition a unit of mass that describes the mass of an atom or molecule atomic number average atomic mass atomic mass unit mass number 1 of 19 term atomic number atoms described as electrons orbiting the nucleus in well defined paths sum of protons and neutrons how many electrons are in the nucleus of an atom with an atomic number of 15 how many neutrons are in the nucleus of an atom with an atomic number of 25 what is the mass number of an atom with 3 protons 4 neutrons and 3 electrons atomic 10 mass 20 of p of n ne of e atomic 12 mass 24 of p of n of e mass 35 of p cl of e atomic 2 mass 4 of p of n he of e atomic 14 mass 28 of p si of e basic atomic structure worksheet answers 1 a protons b neutrons c electrons a positive b neutral c negative 2 atomic number or identity charge 3 protons electrons in a neutral charge atom only same 4 average atomic weight mass 5 mass number nucleus 6 neutrons protons or atomic number mass number 7 2 1 early ideas in atomic theory 2 2 evolution of atomic theory 2 3 atomic structure and symbolism 2 4 chemical formulas key terms key equations summary exercises atomic structure answer key what is an atom an atom is the smallest unit of matter that retains the chemical properties of an element what are the three main subatomic particles in an atom the three main subatomic particles in an atom are protons neutrons and electrons what is the charge and location of a proton atoms cannot be subdivided created or destroyed how is atomic mass calculated it is the average mass of all the isotopes of an element

the atomic number of an element represents number of protons and the number of electrons charge of an electron 1 the mass number of an element is structure of atom is the second chapter in the ncert class 11 chemistry textbook the topics covered under this chapter include subatomic particles thomson s atomic model rutherford s atomic model bohr s model and the quantum mechanical model of the atom top trending quizzes atomic structure quiz questions and answers would you be able to explain atomic structure then take up this quiz with interactive questions and answers and see for yourself an atom structure consists of a positively charged nucleus surrounded and neutralized by answer d 35 45 35 question 5 the correct electronic configuration of potassium is study with quizlet and memorize flashcards containing terms like the fundamental particle of a chemical element according to dalton s theory is which subatomic particle has smallest mass how mant protons electrons and neutrons does iodine 127 have and more science tech math science atomic structure chemistry quiz atomic structure electron configuration oxidation and more take this chemistry quiz to test yourself about atomic structure including nuclear structure and electron configuration mike agliolo getty images by anne marie helmenstine ph d updated on march 06 2017 1 most of the atom is empty space the rest consists of three basic types of subatomic particles protons neutrons and electrons the protons and neutrons form the atom s central nucleus the ordinary hydrogen atom is an exception it contains one proton but no neutrons

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answer atomic number 50 mass number 118 when referring to an atom we simply use the element s name the term sodium refers to the element as well as an atom of sodium

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answer atomic number z the atomic number of an element is equal to the number of protons present inside the nucleus of its atoms since an isolated atom has no net charge on it in neutral atoms the total number of electrons is equal to its atomic number

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atoms cannot be subdivided created or destroyed how is atomic mass calculated it is the average mass of all the isotopes of an element the atomic number of an element represents number of protons and the number of electrons charge of an electron 1 the mass number of an element is

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