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the molar mass of an element or compound is the mass in grams of 1 mole of that substance a property expressed in units of grams per mole g mol see figure 2 5 1 figure 2 5 1 each sample contains 6 022 10<sup>23</sup> atoms 1 00 mol of atoms from left to right top row 65 4g zinc 12 0g carbon 24 3g magnesium and 63 5g copper

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1 1 introduction to chemistry chemistry is a very universal and dynamically changing subject to be confined to a fixed definition it might be better to think of chemistry more as a point of view that places its major focus on the structure and properties of substances particular kinds of matter and especially on the changes they undergo

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radical chemistry begins by defining the term radical it is exactly 100 years since moose gomberg claimed that he had observed a substance containing a trivalent carbon atom i e a carbon centred free radical nowadays simply a carbon centred radical

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