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microsoft word chemsheets a2 1014 thermodynamics booklet docx section 1 recap of as energetics what is enthalpy enthalpy change  $\Delta H$  it is a measure of the heat content of a substance change in heat content at constant pressure standard conditions  $\Delta H^\circ$  100 kpa and a stated temperature

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lecture 16 thermodynamics ii worksheet 1 a 85.2 g copper bar was heated to 221.32 °C and placed in a coffee cup calorimeter containing 425.0 ml of water at 22.55 °C the final temperature of the water was recorded to be 26.15 °C a how much heat was gained by the water b how much heat was lost by the copper c

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the three laws of thermodynamics describe restrictions on the behavior of virtually the entire physical world we can experience everything that is possible or impossible in a physical chemical or biological system is in some way related to these laws

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when the heat for a process is positive there is always an increase in temperature of the system false not for a phase change or chemistry for each of the following note what you would expect for the entropy of the system surroundings and total

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day 1 transforming energy and thermal energy temperature heat homework 3 x readings below next class quiz transforming energy the laws of thermodynamics

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$c_2h_6(g) + 2o_2(g) \rightarrow 2co_2(g) + 3h_2o(l)$  step 2 calculate standard  $\Delta G$  for the reaction using the standard  $\Delta G$  of formation for the reactants and products and the stoichiometric coefficients since diatomic oxygen is already in its standard form it has a  $\Delta G$  f o of 0

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thermodynamics worksheet fill the blanks in the following sentences with the correct thermodynamics term  
the thing we measure when we want to determine the average kinetic energy of random motion in the  
particles of a substance is temperature

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four methods for finding  $\Delta H$  calculate it using average bond enthalpies measure it using a calorimeter  
calculate it using Hess's law calculate it using enthalpies of formation calorimetry the measurement of  
heat transfer calorimeter a device used to determine the amount of heat transferred heat flow cannot be  
measured directly

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tutorial do the following processes produce an increase or a decrease in entropy  $\text{CO}_2(\text{s}) \rightarrow \text{CO}_2(\text{g})$  this  
process increases the entropy or randomness of the system the molecules in a solid system are locked  
up in a repetitive pattern there is a small degree of randomness in solid systems due to the vibrations of  
individual molecules

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thermodynamics questions google classroom when heating a solution a scientist detects a temperature  
increase in the solution during a period of time which of the following statements accurately characterizes  
the solution during this period choose 1 answer the solution is at boiling point a the solution is at boiling  
point

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using the thermodynamic tables calculate  $\Delta H^\circ_{\text{rxn}}$  and  $\Delta G^\circ_{\text{rxn}}$  for the following reactions using  $\Delta H^\circ_f$  and  $\Delta G^\circ_f$  for products and reactants

1  $\Delta H^\circ_{\text{rxn}} = \sum \Delta H^\circ_f(\text{products}) - \sum \Delta H^\circ_f(\text{reactants})$  and  $\Delta G^\circ_{\text{rxn}} = \sum \Delta G^\circ_f(\text{products}) - \sum \Delta G^\circ_f(\text{reactants})$

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using the second law of thermodynamics explain why heat flows from a hot body to a cold body but not from a cold body to a hot body one test of the spontaneity of a reaction is whether the entropy of the universe increases  $\Delta S_{\text{univ}} > 0$  using an entropic argument show that the following reaction is spontaneous at 25 °C

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thermodynamics ii clausius clapeyron equation the clausius clapeyron equation was initially proposed by a german physics rudolf clausius in 1834 and then further developed by french physicist benoit clapeyron in 1850 this equation is extremely

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as section 1 recap of as energetics what is enthalpy enthalpy change  $\Delta H$  it is a measure of the heat content of a substance change in heat content at constant pressure standard conditions  $p^\circ = 100 \text{ kPa}$  and a stated temperature

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as we saw last week enthalpy and internal energy are state functions which means that the sum of the heats of any set of steps that adds to give an overall reaction will have the same heat as doing

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